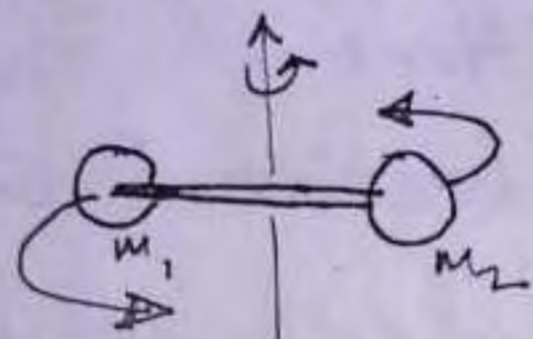


# LINE SPLITTING DUE TO A SIMPLE, RIGID ROTATOR...

(This is a naive analysis)

A "simple, rigid rotator" has only one nonzero value for the moment of inertia,  $I$ , and the value along one axis vanishes. An example is a "dumbbell"



rotation axis, through center of mass.

The kinetic energy due to rotation is then:

$$E_{\text{rot}} = \frac{L^2}{2I}$$

where  $L$  is the angular momentum.

Angular momentum is quantized (total angular momentum)

as:  $L_J^2 = \hbar^2 J(J+1)$ , where  $J = 0, 1, 2, \dots$

$\therefore$  The energy of the  $J^{\text{th}}$  rotational state is:

$$E_J = \frac{\hbar^2}{2I} J(J+1) \quad (1)$$

Dipole emission selection rules allow for only  $\Delta J = \pm 1$  transitions, so:

$$|E_J - E_{J-1}| = \frac{\hbar^2}{2I} [J(J+1) - (J-1)J] = \frac{\hbar^2}{I} J \equiv \Delta E_J \quad (2)$$

(2)

So a simple, rigid rotator has a rotational spectrum consisting of equally-spaced lines with spacing of  $\Delta E_J = \frac{h^2}{I}$ .

In the case of an ~~atomic~~ electron transition, in a simple diatomic molecule, the line will be split because of these molecular rotational transitions. A further splitting is caused by vibrational energy transitions of the molecular bond.

Let's ignore all these complications and assume that an electronic transition gives rise to a visible emission line at energy  $E_0$ , but this line is split into many individual lines due to the various rotational energies & transitions among the different molecules emitting the line.

In this case,  $E_J = E_0 \pm \Delta E_J = E_0 \pm \frac{h^2}{I} J$

Since  $E = h\nu$ , and  $\nu = \frac{c}{\lambda}$

$$\frac{h}{\lambda_J} = \frac{h}{\lambda_0} \pm \frac{h^2}{cI} J \rightarrow \frac{1}{\lambda_J} = \frac{1}{\lambda_0} \pm \frac{h}{2\pi cI} J$$

Two adjacent lines in the spectrum would be separated by:

$$\frac{1}{\lambda_J} - \frac{1}{\lambda_{J-1}} = \frac{h}{2\pi c I} \quad (3)$$

Let's use the above equation to estimate the bond length of a diatomic molecule.

For a molecule with atoms  $m_1$  &  $m_2$  and bond length  $r$ , the moment of inertia about the center of mass is given by:

$$I = \mu r^2, \quad \mu = \frac{m_1 m_2}{m_1 + m_2} \quad (4)$$

$\mu$  is the reduced mass. If we measure the mass using energy units ( $E = mc^2$ )

$$\text{Then } c^2 I = \mu r^2 \quad (\mu \text{ in energy units})$$

Substituting in (3):

$$\frac{1}{\lambda_J} - \frac{1}{\lambda_{J-1}} = \frac{hc}{2\pi \mu r^2} \quad (\mu \text{ in energy units}) \quad (5)$$

(T)  
Measuring  $\lambda$  &  $r$  in Angstroms, with  
 $hc = 1970 \text{ eV}\text{\AA}$ , and measuring  $\mu$  in  
proton masses, we get:

$$\left( \text{since } \frac{1}{\lambda_J} - \frac{1}{\lambda_{J-1}} = \frac{\lambda_{J-1} - \lambda_J}{\lambda_J \lambda_{J-1}} \right)$$

$$\left( \frac{r}{\text{\AA}} \right) \approx \sqrt{\frac{(\lambda_J/\text{\AA})(\lambda_{J-1}/\text{\AA})}{(3.0 \times 10^6) \left( \frac{\mu}{m_p} \right) (\lambda_{J-1}/\text{\AA} - \lambda_J/\text{\AA})}}$$

Question: for OH,  $\mu \approx m_p$ . with  
 $\lambda_J \approx \lambda_{J-1} \approx 3080 \text{\AA}$   
and  $\lambda_{J-1} - \lambda_J \approx 1.8 \text{\AA}$

what would  $r$  be?